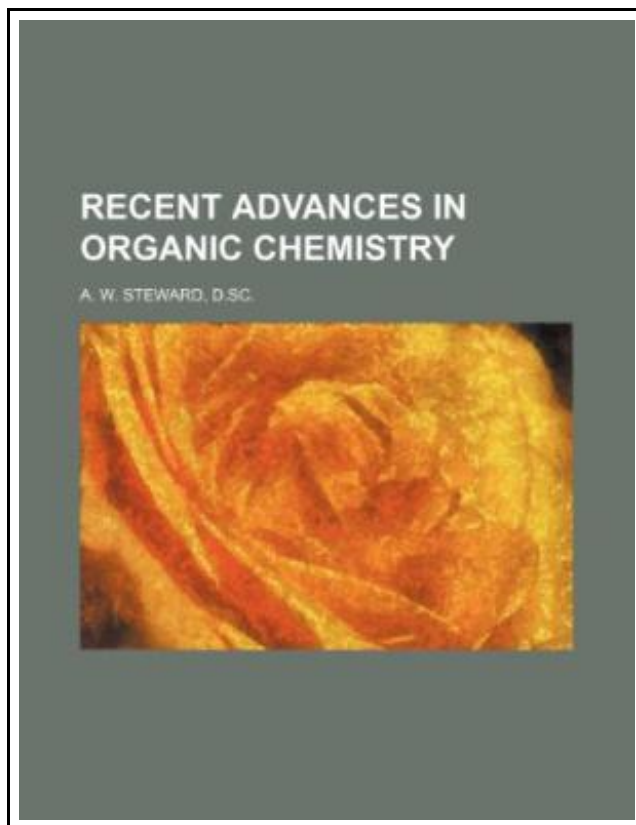


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Reviews

This book will not be simple to get going on reading but extremely exciting to read through. Yes, it can be play, still an interesting and amazing literature. I am very easily could possibly get a delight of reading a written book.

(Rene Olson)

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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1908 Excerpt: . $\text{CH}_3 / \text{CH} \text{CH}_2\text{--C} \text{CH}_2 \text{I} / \text{I} \text{OH} \text{CH--CO} \text{I} \text{CH}_3$ When thujone is reduced it yields thujyl alcohol; and from this we can produce the thujyl derivative of xanthogenic acid (sulphothio-carbonic acid). When this substance is distilled it 1 Semmler, Ber., 1900, 88, 275, 24.59. breaks down into carbon oxysulphide, methyl mercaptan, and a hydrocarbon thujene-- $\text{C}_{10}\text{H}_{17}\text{O}$. $\text{CS} \cdot \text{SCH}_3 = \text{COS} + \text{CH}_3\text{SH} + \text{C}_{10}\text{H}_{16}$ Tschugaeffl has shown that thujene thus obtained is a mixture of two hydrocarbons, to which he attributes the formulae-- $\text{C}_3\text{H}_7 \text{C}_3\text{H}_7 \text{I} \text{I} \text{CH}_2\text{--C--CH}_2 \text{CH}_2\text{--C--CH} \text{I} / \text{I} \text{I} / \text{II} \text{CH--C=CH} \text{CH--C--CH} \text{CH}_3 \text{CH}_3$ a-Tlmjene 6-Thujene. Sabinene is a terpene isomeric with thujene. The corresponding alcohol, sabinol, gives a trihydric alcohol on oxidation with permanganate, so that it probably contains a double bond. On reduction, sabinol gives thujyl alcohol. With alcoholic sulphuric acid it is transformed into cymene, while with zinc dust it is converted into thujone. From these data it is evident that a close relation exists between sabinol and thujone, and that probably they contain the same skeleton of carbon atoms. Now, when the trihydric alcohol obtained by oxidizing sabinol is treated with a trace of acid it loses two molecules of water and is converted into cumyl alcohol-- $\text{C}_{10}\text{H}_{13}\text{O}_3 = 2\text{H}_2\text{O} + \text{C}_3\text{H}_7 \cdot \text{C}_6\text{H}_4 \cdot \text{CH}_2\text{OH}$ This proves that the double bond in sabinol must lie outside the ring, as otherwise the group-- CH_2OH could not be formed. Taking all these properties into account, and assuming the correctness of...



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