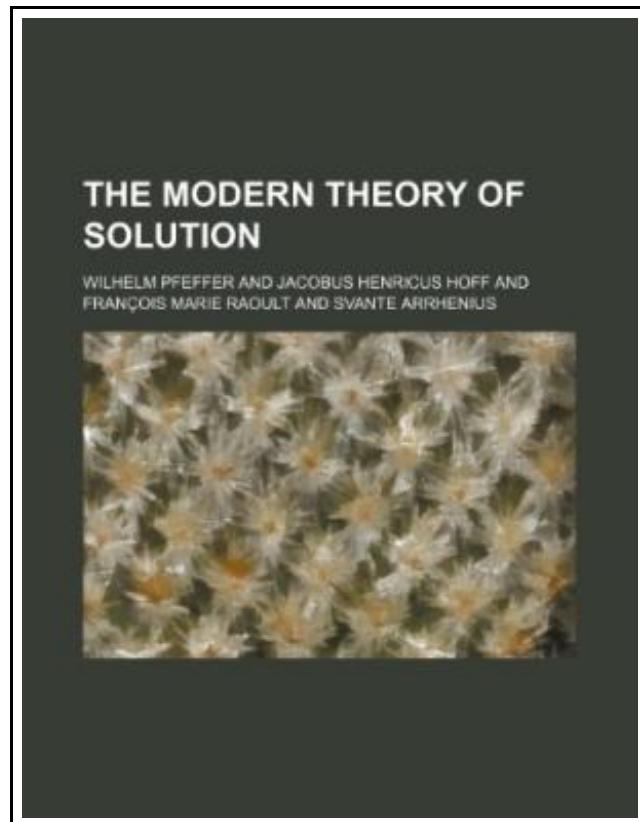


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1899 Excerpt: . of the existence of molecular complexes, f But, as stated, this can also be ascribed to the conversion of inactive into active molecules. Further, some chemists, to support the idea of constant valence, would assume J molecular complexes, in which the unsaturated bonds could become saturated. But the doctrine of constant valence is so much disputed that we are scarcely justified in basing any conclusions upon it. The conclusions thus arrived at, that, e.g., potassium chloride would have the formula (KCT)3, L. Meyer Hittorf, Pogg. Ann., 106, 547 and 551 (1859); Wied. Etektr., 2, 584. f Ostwald, Lebrb. all. Chem., I., 811; L. Meyer, Moderne Theorien der Chemie, p. 319 (1880). h. Meyer, I. e. p. 360. sought to support by the fact that potassium chloride is much less volatile than mercuric chloride, although the former has a much smaller molecular weight than the latter. Independent of the theoretical weakness of such an argument, this conclusion could, of course, hold only for the pure substances, not for solutions. Several other reasons have been brought forward by L. Meyer for the existence of molecular complexes, e.g., the fact that sodium chloride diffuses more slowly than hydrochloric acid, but this is probably to be referred to the greater friction (according to electrical determinations), of sodium against water, than of hydrogen. But it suffices to cite L. Meyer s own words: Although all of these different points of departure for ascertaining molecular weights in the liquid condition are...

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